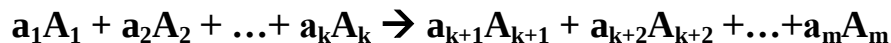


STOICHIOMETRY

Three types:

Molar:

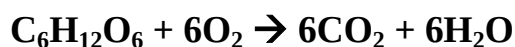


where A_i are reactants for $i = 1 \rightarrow k$ and products for $i = k+1 \rightarrow m$

And a_i are stoichiometric coefficients

Characteristics of molar stoichiometric relations:

Example:

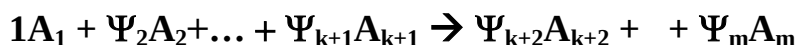


Units: moles, molecules

Element mass is conserved

Charge is conserved

Mass:



Where Ψ_i are stoichiometric coefficients

And $\Psi_i = \frac{a_i(\text{MW of } A_i)}{a_1(\text{MW of } A_1)}$ for $i = 2 \rightarrow m$

Note that mass stoichiometric coefficients are normalized by the molecular weight of the reference compound (typically reactant 1).

Units: mass (grams, kilograms)

Charge and element mass is not conserved

Total mass of reactants = total mass of products

That is:

$$\sum_{i=1}^k \Psi_i = \sum_{i=k+1}^m \Psi_i$$

Example:



Reference compound is glucose, MW = 180

$$\Psi_1 = 1$$

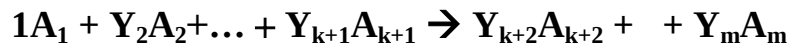
$$\Psi_2 = 6(32)/180 = 1.067 \text{ g-O}_2/\text{g-glucose}$$

$$\Psi_3 = 6(44)/180 = 1.467 \text{ g-CO}_2/\text{g-glucose}$$

$$\Psi_4 = 6(18)/180 = 0.60 \text{ g-H}_2\text{O/g-glucose}$$

$$\begin{aligned} \text{balance:} \quad & 1 + 1.067 = 2.067 \\ & 1.467 + 0.6 = 2.067 \end{aligned}$$

COD (electron):



where the reference compound is COD and the stoichiometric coefficients, Y_i are in units of mass COD

$$Y_1 = 1$$

$$Y_i = \frac{a_i(MW - A_i)(COD_i)}{a_1(MW - A_1)(COD_1)} = \Psi_i \frac{COD_i}{COD_1}$$

Where COD_i is the COD equivalent of the *i*th compound. For example the COD equivalent of glucose is $6 \cdot 32/180 = 1.067 \text{ g-COD/g-glucose}$ (see molar stoichiometry).

Other COD equivalents: $COD_{\text{cells}}(C_5H_7NO_2) = 5(32)/113 = 1.42 \text{ g-COD/g-cells}$

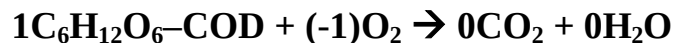
$$COD_{O_2} = -1 \text{ g-COD/g-O}_2$$

COD equivalent of water, CO₂, and any other compound not involved in exchange of electrons is zero.

Units: g-COD

COD mass is conserved

Example:



$$\begin{aligned} \text{Balance:} \quad & 1 - 1 = 0 \\ & 0 + 0 = 0 \end{aligned}$$